

Standard: #2-2

Chemistry

Electron Configurations and Orbital Notations

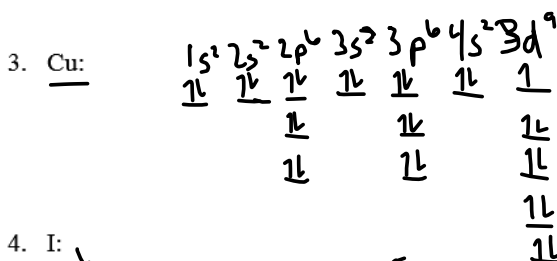
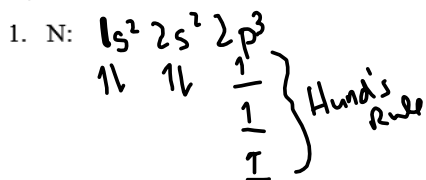
Objectives:

What is the electronic structure of an atom?

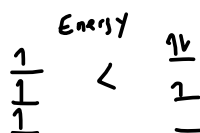
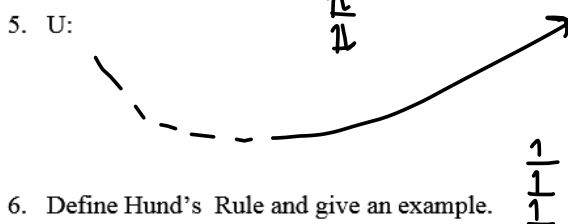
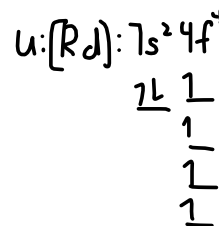
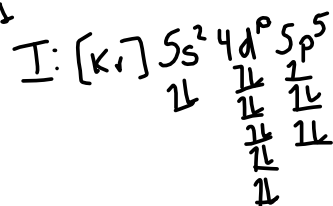
- Student will be able to write electron configurations.
- Student will be able to write orbital diagrams.



For each of the following provide the Long and short hand Electron Configurations and the orbital notations.



Short → 4. I: ↓



6. Define Hund's Rule and give an example.

7. Define the Pauli exclusion principle (Use the internet)

No $2e^-$ in an atom can have same quantum #s

